

THE ARCHAEOLOGY OF TOBACCO PIPES
IN EASTERN NORTH AMERICA

SMOKING and CULTURE

Edited by Sean M. Rafferty and Rob Mann

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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c is the limiting reagent 4.33 g of h 2 are left over 13

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 $80 \text{ g I}_2\text{O}_5 \times \frac{1 \text{ mol I}_2\text{O}_5}{1 \text{ mol I}_2} \times \frac{1 \text{ mol I}_2}{1 \text{ mol I}_2\text{O}_5} \times \frac{333.8 \text{ g I}_2}{1 \text{ mol I}_2} = 333.8 \text{ g I}_2$

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 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$