

Acid Base Strength Table

TABLE 8.2 Some Acids and Their Conjugate Bases, in Decreasing Order of Acid Strength

Acid		Conjugate Base			
<div>Strong Acids</div> <div></div> <div>Weak Acids</div>	HI	Hydroiodic acid	I ⁻	Iodide ion	
	H ₂ SO ₄	Sulfuric acid	HSO ₄ ⁻	Hydrogen sulfate ion	
	HCl	Hydrochloric acid	Cl ⁻	Chloride ion	
	HNO ₃	Nitric acid	NO ₃ ⁻	Nitrate ion	
	H ₃ O ⁺	Hydronium ion	H ₂ O	Water	
	HSO ₄ ⁻	Hydrogen sulfate ion	SO ₄ ²⁻	Sulfate ion	
	H ₃ PO ₄	Phosphoric acid	H ₂ PO ₄ ⁻	Dihydrogen phosphate ion	
	CH ₃ COOH	Acetic acid	CH ₃ COO ⁻	Acetate ion	
	H ₂ CO ₃	Carbonic acid	HCO ₃ ⁻	Bicarbonate ion	
	H ₂ S	Hydrogen sulfide	HS ⁻	Hydrogen sulfide ion	
	(H ₂ PO ₄) ⁻	Dihydrogen phosphate ion	HPO ₄ ²⁻	Hydrogen phosphate ion	
	NH ₄ ⁺	Ammonium ion	NH ₃	Ammonia	
	C ₆ H ₅ OH	Phenol	C ₆ H ₅ O ⁻	Phenoxide ion	
	HCO ₃ ⁻	Bicarbonate ion	CO ₃ ²⁻	Carbonate ion	
	HPO ₄ ²⁻	Hydrogen phosphate ion	PO ₄ ³⁻	Phosphate ion	
	H ₂ O	Water	OH ⁻	Hydroxide ion	
	C ₂ H ₅ OH	Ethanol	C ₂ H ₅ O ⁻	Ethoxide ion	
	NH ₃	Ammonia	NH ₂ ⁻	Amide ion	
					<div>Weak Bases</div> <div></div> <div>Strong Bases</div>

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Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Mike Jess' text.

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Solid Acids and Bases Kozo Tanabe,2012-12-02 Solid Acids and Bases Their Catalytic Properties reviews developments in the studies of acidic and basic properties of solids including the efficacy and special characteristics of solid acid and base catalysts This book discusses the determination of basic and acidic properties on solid surfaces and relationship between acid strength and acid amount The structure and acid base properties of mixed metal oxides and correlation between acid base properties and catalytic activity and selectivity are also deliberated This publication is useful to professional chemists and graduate students in the fields of organic inorganic and physical chemistry petroleum chemistry and catalysis including readers interested in the acidic and basic properties on solid surfaces

More Acids and Bases ,1944 **Acids and Bases** Brian G. Cox,2013-01-31 Acids and bases are ubiquitous in chemistry Our understanding of them however is dominated by their behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents

and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described

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